

1. Heat the mixture: $2.26 \text{ kJ} = 50^\circ\text{C} \times 1.00 \text{ g} + 0.50 \text{ g} \times 4.18 \text{ J/g}^\circ\text{C}$
 2. How much water is formed? $9.00 \text{ mol CO}_2 = \frac{9.00 \text{ mol CO}_2 \times 44.01 \text{ g/mol}}{44.01 \text{ g/mol}} = 450.9 \text{ g H}_2\text{O}$
 3. How much heat is released during the formation of this water? $9.00 \text{ mol H}_2\text{O} \times \frac{285.8 \text{ kJ}}{1 \text{ mol H}_2\text{O}} = 2572.2 \text{ kJ}$
 4. How much heat is released during the formation of this water? $9.00 \text{ mol H}_2\text{O} \times \frac{285.8 \text{ kJ}}{1 \text{ mol H}_2\text{O}} = 2572.2 \text{ kJ}$
 5. If 2.26 kJ of heat are produced how much water is released? $2.26 \text{ kJ} = \frac{2.26 \text{ kJ}}{2572.2 \text{ kJ}} \times 450.9 \text{ g H}_2\text{O} = 0.39 \text{ g H}_2\text{O}$
 6. Calculate the amount of heat (in kJ) released by 200 g of water from 20°C to 5°C. $200 \text{ g} \times 4.18 \text{ J/g}^\circ\text{C} \times (5^\circ\text{C} - 20^\circ\text{C}) = -14652 \text{ J} = -14.65 \text{ kJ}$
 7. 20.0 kJ of heat are added to a 200 g sample of water initially at 10°C. Calculate the temperature of the water sample. $200 \text{ g} \times 4.18 \text{ J/g}^\circ\text{C} \times (\Delta T) = 20000 \text{ J}$
 $\Delta T = \frac{20000 \text{ J}}{836 \text{ J/}^\circ\text{C}} = 23.9^\circ\text{C}$
 $T_{\text{final}} = 10^\circ\text{C} + 23.9^\circ\text{C} = 33.9^\circ\text{C}$

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Stoichiometry Stoichiometric Calculations Starting with 10. g of C₆H₁₂O₆... we calculate the moles of C₆H₁₂O₆... use the coefficients to find the moles of H₂O & CO₂ and then turn the moles to grams C₆H₁₂O₆(s) + 6 O₂(g) → 6 CO₂(g) + 6 H₂O(l) 10.g ? + ? Example: 10 grams of glucose (C₆H₁₂O₆) react in a combustion reaction. How many grams of each product are

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8.2 Stoichiometric calculations (ESBP8) In grade 10 you learnt how to write balanced chemical equations and started looking at stoichiometric calculations. By knowing the ratios of substances in a reaction, it is possible to use stoichiometry to calculate the amount of either reactants or products that are involved in the reaction.

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