

Name: Kur Period: _____ Date: _____

Stoichiometry Introduction
Accelerated Chemistry

1. 2 buns + 1 patty → 1 hamburger (1 bun = 50g, 1 patty = 500g) if you have 2,000g of buns and 80,000g of patties, how many hamburgers can you make?

$$2000g \text{ buns} \times \frac{1 \text{ bun}}{50g} \times \frac{1 \text{ hamburger}}{2 \text{ buns}} = 2 \text{ hamburgers}$$

$$80000g \text{ patties} \times \frac{1 \text{ patty}}{500g} \times \frac{1 \text{ hamburger}}{1 \text{ patty}} = 160 \text{ hamburgers}$$

2. 5 beads + 2 strings → 1 bracelet (1 bead = 10g, 1 string = 20g) if you have 500g of beads and 800g of string, how many bracelets can you make?

$$500g \text{ beads} \times \frac{1 \text{ bead}}{10g} \times \frac{1 \text{ bracelet}}{5 \text{ beads}} = 10 \text{ bracelets}$$

$$800g \text{ string} \times \frac{1 \text{ string}}{20g} \times \frac{1 \text{ bracelet}}{2 \text{ strings}} = 20 \text{ bracelets}$$

3. 3 diamonds + 1 ring → 1 engagement ring (1 diamond = 2 carats, 1 ring = 200g platinum) if you have 30 carats of diamonds and 600g of platinum, how many rings can you make?

$$30 \text{ car} \times \frac{1 \text{ dia}}{2 \text{ car}} \times \frac{1 \text{ ER}}{3 \text{ dia}} = 5 \text{ ER}$$

$$600g \text{ Pt} \times \frac{1 \text{ ring}}{200g \text{ Pt}} \times \frac{1 \text{ ER}}{1 \text{ ring}} = 3 \text{ ER}$$

4. Design a generalized set of instructions to solve problems like the ones you see above

- ① Convert known quantities into unknown quantities using given information
- ② # made = least amount

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1. a. 2 / 13 b. 13 / 8 c. 13 / 10 d. 2 / 8 (or 1 / 4) e. 2 / 10 (or 1 / 5) 2. The KClO₃ / O₂ molar ratio is 2/3. 2 mol KClO₃ / 3 mol. O₂ = 12.00 mol KClO₃ / x = 18.00 mol. x = 18.00 mol of O₂ 3. Given the following equation: 2 K + Cl₂ ---> 2 KCl How many grams of KCl is produced from 2.50 g of K and excess Cl₂ ...

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W/ answers Website Upload Big Numbers and Chemistry At the most fundamental level, the chemist needs a unit that describes a very large quantity. One of the most well-known numbers in the study of chemistry is number of units in a mole. The number of **Unit 6: Reactions and Stoichiometry** moles amp stoichiometry. packet answer key 45 55 180 207. mole homework packet answer key name 1 i ' hour. mole concept homework packet answer key gaywix de. classwork and homework handouts. passat manual pdf fraggorillazz de. mole concept homework packet answer key celsen de. brilliant coad3 pdf http brilliant coad3 org. moles amp stoichiometry answers key questions amp exercises. mole ...

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Stoichiometry: Part 2 Mass Mole Relationshi Block: To convert nncs to mass we use the mass of one mole. moles of R molar mass of B 1 mole of B For example: How many grams of ethane are in 13.3 moles of ethane? Because wc know that thetQS 30.07 g of ethane in one mole: moles g of mole C, Hs motes use the of mote of A Of A Ot A Of ethane Ot know that g Ot ethane in one moles t: totlowing to ...

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Then, multiply it by the number of moles to get an answer in grams: grams of sample = (molar mass) x (moles) For example, find the number of grams in 0.700 moles of hydrogen peroxide, H₂ O₂. Calculate the molar mass by multiplying the number of atoms of each element in the compound (its subscript) times the atomic mass of the element from the periodic table. Molar mass = (2 x 1.008) + (2 x ...

How to Convert Grams to Moles and Vice Versa

You can't use it the same way that you do in English when you answer, so maybe its the multiple duties that make "do" the linguistic outlier. #8:: Doug:: (view all by):: June 12, 2017, 01:49 AM: And from near the end, "In Old English, however, 'Ving-Thor was mad when he woke up' would have been Wraþmod wæs Ving-bórr/he áwæcnede. We can just about wrap our heads around this as ...

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