Name _____ Date_____ Period _____

Solutions Worksheet #1: Molarity

1) Complete the table:

	Solute formula	Molar Mass of solute (g/mole)	Mass of Solute (g)	Moles of solute (mole)	Molarity of solution (M)	Volume of solution (L)
a	NaCl	58.5	58.5	1	1.0	1.0
b	NaOH		117.0			4.0
c	$MgCl_2$	95.3	190.3	2	1.0	2.0
d	NaCl		292.5		0.5	
e	KBr	119	238.0	2	2.0	1.0
f	NaCl				1.5	1.0
g	NaCl	58.5	14.63	0.25	1.0	0.25
h	$FeCl_3$				0.1	0.50
i	HCl	36.5	438.0	12	6.0	2.0
j	NH_3		39.1		3.0	

2) Calculate the number of moles and the number of grams of solute in each solution:

		Moles of solute (mol)	Molar Mass of solute (g/mol)	Mass of solute (g)
a	1.0 liter of .5 M NaCl	.5	58.5	29.3
b	500.0 mL of 2.0 M KNO_3			
c	250.0 mL of .10 M $CaCl_2$	0.25	111	2.78
d	2.0 liters of .30 M Na_2SO_4			

3) Compute the mass of solute needed to make 500.0 mL of solution at the indicated molarity.
a: 0.5M sulfuric acid 24.5g
b: 0.01M HNO_3
c: 6.0M HCl 109.5g
d: 0.50M sodium hydroxide
e: 0.1M iron III chloride 8.125g
f: 3M NH_3
g: 5.0M KOH 140.0g

4) How many moles of H_2SO_4 are in 1.00 liter of a 1.55M H_2SO_4 solution? How many grams?

5) How many grams of sodium sulfate are contained in 1.50L of .25M solution? 53.25g Na_2SO_4

6) How many grams of ammonium sulfate are required to prepare 3.50L of a 1.55M solution?

7) How many moles of sodium chromate are contained in 1.75 L of a 2.00M solution?
3.5 mol Na_2CrO_4

8) A sample of glucose ($C_6H_{12}O_6$) is dissolved in water. How many moles of glucose are dissolved in 200.0 ml of solution if its concentration is .150M?

9) A mass of 98g of sulfuric acid are dissolved in water to prepare a .50M solution. What is the volume of the solution, in liters? 2L

10) What is the molarity of a solution of HNO_3 that contains 12.6g of solute in 500ml of solution?

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Solutions Worksheet 1 Molarity Answer Key

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Molarity Worksheet # 1
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Molarity Worksheet W 331 Everett Community College Student Support Services Program What is the molarity of the following solutions given that: 1) 1.0 moles of potassium fluoride is dissolved to make 0.10 L of solution. 2) 1.0 grams of potassium fluoride is dissolved to make 0.10 L of solution.

Molarity Worksheet W 331 - Everett Community College
Chemistry: Molarity of Solutions Directions: Solve each of the following problems. Show your work and include units for full credit. 1. What mass of the following chemicals is needed to make the solutions indicated? a. 1.0 liter of a 1.0 M mercury (II) chloride (HgCl2) solution. b. 2.0 liters of a 1.5 M sodium nitrate (NaNO3) solution. c. 5.0 liters of a 0.1 M Ca(OH)2 solution. d. 100 mL of a ...

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1 liter solution. What is the molarity of a solution that contains 4.53 moles of lithium nitrate in 2.85 liters of solution? 4.53 mol LiNO3 = 1.59 M LiN03. 2.85 L soln . What is the molarity of a solution that contains 0.00372 moles hydrochloric acid in 2.39 x 10-2 liters of solution? 0.00372 mol HCL = 0.156 M HCL. 2.39x10-2 L soln. A flask contains 85.5 g C12H22O11 (sucrose) in 1.00 liters of ...

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5. 125 cm 3 of solution contains 3.5 moles of solute. What is the molarity of the solution? ? g KNO 3 = 0.175 mol KNO 3 × 101.1 g KNO 3 1 mol KNO 3 = 17.7 g KNO 3 M = 3.5 mol 0.125 L = 28 M 6. Which solution is more concentrated? Solution “A” contains 50.0 g of CaCO3 in 500.0 mL of solution. Solution “B” contains 6.0 moles of H 2SO 4 ...

Molarity: Molarity = 1. 2. - Central Bucks School District
Molarity Worksheet # 2 identifiera _____ What does molarity mean? Number of moles of solute. 1 liter solution. What is the molarity of a solution that contains 4.53 moles of lithium nitrate in 2.85 liters of solution? 4.53 mol LiNO 3 = 1.59 M LiN0 3. 2.85 L soln

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